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consider the following mechanism
 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$ fast
 $\text{O} + \text{O}_2 \rightarrow \text{O}_3$ slow
a) write the overall balanced chemical equation
b) identify any intermediates within the mechanism

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 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
g the reaction between Mg and HCl is represented by the equation above in a kinetics experiment
a) 0.080 g sample of Mg ribbon is added to an Erlenmeyer flask containing 50 mL of 3.0 M HCl at 25 °C

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