Chapter 2:

The Earth's atmosphere

PROBLEMS/SOLUTIONS

 The mixing ratio of oxygen in the atmosphere is 20.95%. Calculate the concentration in mol L⁻¹ and in g m⁻¹ at P* (101 325 Pa, 1.00 atm) and 25°C.

Sedution

The mixing ratio for O₂ (g) is 20.95% (Table 2.1). Use the 'Ideal Gas' law (PV= nRT) to calculate the total number of moles of gas in 1.00 L for the given conditions. Constants are given in Appendices B.1 and C.1.

 $T = 25^{\circ}C$ (298 K), $P = 1.01325 \times 10^{5}$ Pa. R = 8.315 J K⁻¹ mol⁻¹, and V = 1.00 L = 1.00×10^{-3} m³

The total number of moles of gas in 1.00 L is

n = PV/RT n = 0.04089 mol

The number of moles of oxygen can be calculated from the mole fraction (which is the same as the mixing ratio).

% $O_2 = 20.95$, mole fraction = 0.2095

The number of moles of O_2 in 1.00 L is: 0.2095 x 0.04089 = 0.00857 mol

The concentration of O2 (g) is: 8.57 x 10⁻³ mol L⁻¹.

The concentration of O2 (g) in units of g m3 is determined as follows:

1 m3 = 1000 L, therefore, 1000 L contains 8.57 moles of O2 (g). (M.M. of O2 is 31.9988 g mol 5)

 $31.9988 \text{ g mol}^{-1} \times 8.57 \text{ mol m}^{-3} = 274 \text{ g m}^{-3}$

The O₂ (g) concentration can be expressed as either 8.57 x 10⁻³ mol L⁻¹ or 274 g m⁻².

2. Calculate the atmospheric pressure at the stratopause. What are the concentrations (mol m⁻¹) of dioxygen and dinitrogen at this altitude? How do these concentrations compare with the corresponding values at sea level?

Solution

Use Equation 2.3 to calculate the pressure at 50 km (stratopause).

$$P_{s} = P^{\circ}e^{-\overline{M}_{s} \cdot g}$$

 $R = 8.315 \text{ J K}^{-1} \text{ mol}^{-1}$, $P^* = 101.325 \text{ Pa}$, $g = 9.81 \text{ m s}^{-2}$, $M_u = 0.02896 \text{ kg mol}^{-1}$ (average molar mass of air), h = 50 km (50.000 m) and $T = -2^{\circ}\text{C}$ (271 K). Note: these last two parameters are the approximate altitude and temperature at the stratopause.

Solutions Manual For Environmental Chemistry

Colin Baird, Michael Cann

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