

Standard Reduction Potentials in Aqueous Solution at 25° C

Half-reaction			$E^\circ(\text{V})$
$\text{F}_2(\text{g}) + 2 e^-$	$\rightarrow$	2F^-	2.87
$\text{Co}^{3+} + e^-$	$\rightarrow$	Co^{2+}	1.82
$\text{Au}^{3+} + 3 e^-$	$\rightarrow$	$\text{Au}(\text{s})$	1.50
$\text{Cl}_2(\text{g}) + 2 e^-$	$\rightarrow$	2Cl^-	1.36
$\text{O}_2(\text{g}) + 4 \text{H}^+ + 4 e^-$	$\rightarrow$	$2 \text{H}_2\text{O}(\text{l})$	1.23
$\text{Br}_2(\text{l}) + 2 e^-$	$\rightarrow$	2Br^-	1.07
$2 \text{Hg}^{2+} + 2 e^-$	$\rightarrow$	Hg_2^{2+}	0.92
$\text{Hg}^{2+} + 2 e^-$	$\rightarrow$	$\text{Hg}(\text{l})$	0.85
$\text{Ag}^+ + e^-$	$\rightarrow$	$\text{Ag}(\text{s})$	0.80
$\text{Hg}_2^{2+} + 2 e^-$	$\rightarrow$	$2 \text{Hg}(\text{l})$	0.79
$\text{Fe}^{3+} + e^-$	$\rightarrow$	Fe^{2+}	0.77
$\text{I}_2(\text{s}) + 2 e^-$	$\rightarrow$	2I^-	0.53
$\text{Cu}^+ + e^-$	$\rightarrow$	$\text{Cu}(\text{s})$	0.52
$\text{Cu}^{2+} + 2 e^-$	$\rightarrow$	$\text{Cu}(\text{s})$	0.34
$\text{Cu}^{2+} + e^-$	$\rightarrow$	Cu^+	0.15
$\text{Sn}^{4+} + 2 e^-$	$\rightarrow$	Sn^{2+}	0.15

Standard Potentials In Aqueous Solution

Cynthia G. Zoski



Standard Potentials In Aqueous Solution:

Standard Potentials in Aqueous Solution Allen J. Bard, Roger Parsons, Joseph Jordan, 1985-08-27 The best available collection of thermodynamic data The first of its kind in over thirty years this up to date book presents the current knowledge on Standard Potentials in Aqueous Solution Written by leading international experts and initiated by the IUPAC Commissions on Electrochemistry and Electroanalytical Chemistry this remarkable work begins with a thorough review of basic concepts and methods for determining standard electrode potentials Building upon this solid foundation this convenient source proceeds to discuss the various redox couples for every known element The chapters of this practical time saving guide are organized in order of the groups of elements on the periodic table for easy reference to vital material AND each chapter also contains the fundamental chemistry of elements numerous equations of chemical reactions easy to read tables of thermodynamic data and useful oxidation state diagrams Standard Potentials in Aqueous Solution is an ideal handy reference for analytical and physical chemists electrochemists electroanalytical chemists chemical engineers biochemists inorganic and organic chemists and spectroscopists needing information on reactions and thermodynamic data in inorganic chemistry And it is a valuable supplementary text for undergraduate and graduate level chemistry students

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Butler,1998-04-13 A celebrated classic in the field updated and expanded to include the latest computerized calculation techniques In 1964 James N Butler published a book in which he presented some simple graphical methods of performing acid base solubility and complex formation equilibrium calculations Today both the book and these methods have become standard for generations of students and professionals in fields ranging from environmental science to analytical chemistry Named a Citation Classic by the Science Citation Index in 1990 the book *Ionic Equilibrium* continues to be one of the most widely used texts on the subject So why tamper with near perfection by attempting a revision of that classic The reason is simple the recent rapid development and wide availability of personal computers In the revised *Ionic Equilibrium* Dr Butler updates his 1964 work by abandoning the slide rule and graph paper for the PC spreadsheet He also expands the original coverage with extensive material on basic principles and recent research The first part of *Ionic Equilibrium* is devoted to the fundamentals of acid base solubility and complex formation equilibria In the second part the author discusses oxidation reduction equilibria develops the principles of carbon dioxide equilibria presents case studies demonstrating the ways in which carbon dioxide equilibria are used in physiology and oceanography and explores the possibility of a pH scale for brines The concluding chapter written by David R Cogley gives examples of general computer programs that are capable of performing equilibrium calculations on systems of many components Replete with real world examples details of important calculations and practical problems *Ionic Equilibrium* is an ideal course text for students of environmental chemistry engineering or health analytical chemistry oceanography geochemistry biochemistry physical chemistry and clinical chemistry It is also a valuable working resource for professionals in those fields as well as industrial chemists involved with solution chemistry *Corrosion and Surface Chemistry of Metals* Dieter Landolt,2007-05-02 Providing a carefully developed and comprehensive overview of the corrosion chemistry of metallic materials this book covers the principal methods of corrosion prevention It includes a systematic study of the physical chemistry of the surface supported by state of the art analysis methods The author builds a scientific foundation by developing t

Handbook of Electrochemistry Cynthia G. Zoski,2007-02-07 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The *Handbook of Electrochemistry* serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory

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The present Supplement Volume Beryllium A 3 continues and completes the description of the physical properties of the element begun in Supplement Volume A 2 1991 and also treats the electrochemical behavior of the metal The unique combination of the Be properties which was pointed out in Supplement Volume A 2 is also demonstrated in the following chapters of this Volume A 3 13 Electrical Properties 14 Electronic Properties 15 Optical Properties Emission and Impact Phenomena 16 Electrochemical Behavior Starting with the electrical properties Be is a rather good electrical conductor in contrast to what might be expected Superconductivity was studied especially on films Quantum effects which are more pronounced in Be than in most other metals are the reason for numerous investigations of the magnetoresistance and the magnetic breakdown effect The basis for many of the characteristic properties is the unique nature of bonding in Be as a consequence of its peculiar electronic structure and the special shape of its Fermi surface which also gave rise to further numerous studies Detailed cluster calculations were performed to better understand the bonding in the metal Regarding the optical properties the high reflectivity of Be particularly in the infrared region makes it attractive for the fabrication of precision optical surfaces mirrors it is also useful for solar collector surfaces in spacecraft applications Emission and electron and ion impact phenomena as well as neutron optics are also discussed *Inorganic Chemistry* Gary Wulfsberg, 2000-03-16

Both elementary inorganic reaction chemistry and more advanced inorganic theories are presented in this one textbook while showing the relationships between the two **Quanta, Matter, and Change** Peter Atkins, Julio de Paula, Ronald Friedman, 2009 aspects of the learning process are fully supported including the understanding of terminology notation mathematical concepts and the application of physical chemistry to other branches of science Building on the heritage of the world renowned Atkins Physical Chemistry *Quanta Matter and Change* gives a refreshing new insight into the familiar by

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