

INORGANIC & ORGANIC CHEM LEC

Ms. Jazheerine Portugal | Chemical Equilibrium

CHEMICAL EQUILIBRIUM

Chemical equilibrium is achieved when the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant.



closed system in equilibrium with its reactants in a closed system at their temperature.



N_2O_4 and 2NO_2 gases at equilibrium.



EQUILIBRIUM CONSTANT

Table shows some experimental data for the reaction just described at 25°C. The gas concentrations are expressed in molarity, which can be calculated from the number of moles of gasses present initially and at equilibrium and the volume of the flask in liters.



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]} = 4.63 \times 10^{-3}$$

Initial Concentrations (M)		Equilibrium Concentrations (M)		Ratio of Concentrations at Equilibrium	
$[\text{NO}_2]$	$[\text{N}_2\text{O}_4]$	$[\text{NO}_2]$	$[\text{N}_2\text{O}_4]$	$\frac{[\text{NO}_2]}{[\text{N}_2\text{O}_4]}$	$\frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$
0.000	0.473	0.0347	0.443	0.00078	4.63×10^{-3}
0.0093	0.464	0.0497	0.404	0.123	4.64×10^{-3}
0.0193	0.450	0.0475	0.391	0.0007	4.60×10^{-3}
0.0400	0.430	0.0923	0.394	0.0009	4.60×10^{-3}
0.200	0.400	0.0784	0.3704	0.0017	4.68×10^{-3}

We can generalize this phenomenon with the following reaction at equilibrium:



For a reversible reaction at equilibrium and a constant temperature, a certain ratio of reactant and product concentrations has a constant value, K (the **equilibrium constant**).

$$K = \frac{[\text{C}]^c[\text{D}]^d}{[\text{A}]^a[\text{B}]^b}$$

HOMOGENEOUS EQUILIBRIA

Homogeneous equilibrium applies to reactions in which all reacting species are in the same phase.



$$K_c = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$K_p = \frac{P_{\text{NO}_2}^2}{P_{\text{N}_2\text{O}_4}}$$

Sample Questions:

- (a) $\text{HF}(\text{aq}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{F}^-(\text{aq})$
 (b) $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$
 (c) $\text{CH}_3\text{COOH}(\text{aq}) + \text{C}_2\text{H}_5\text{OH}(\text{aq}) \rightleftharpoons \text{CH}_3\text{COOC}_2\text{H}_5(\text{aq})$

(a)

$$K_c = \frac{[\text{H}_3\text{O}^+][\text{F}^-]}{[\text{HF}]}$$

(b)

$$K_c = \frac{[\text{NO}_2]^2}{[\text{NO}]^2[\text{O}_2]} \quad K_p = \frac{P_{\text{NO}_2}^2}{P_{\text{NO}}^2 P_{\text{O}_2}}$$

Ssm T A Des Inorganic Chem 4e

**Francesca Deganello, Enrico
Traversa, Ana C. Tavares**



Ssm T A Des Inorganic Chem 4e:

Handbook Of Porphyrin Science: With Applications To Chemistry, Physics, Materials Science, Engineering, Biology And Medicine (Volumes 21-25) Roger Guilard, Kevin M Smith, Karl M Kadish, 2012-06-25 This is the fifth set of Handbook of Porphyrin Science Porphyrins phthalocyanines and their numerous analogues and derivatives are materials of tremendous importance in chemistry materials science physics biology and medicine They are the red color in blood heme and the green in leaves chlorophyll they are also excellent ligands that can coordinate with almost every metal in the Periodic Table Grounded in natural systems porphyrins are incredibly versatile and can be modified in many ways each new modification yields derivatives demonstrating new chemistry physics and biology with a vast array of medicinal and technical applications As porphyrins are currently employed as platforms for study of theoretical principles and applications in a wide variety of fields the Handbook of Porphyrin Science represents a timely ongoing series dealing in detail with the synthesis chemistry physicochemical and medical properties and applications of polypyrrole macrocycles Professors Karl Kadish Kevin Smith and Roger Guilard are internationally recognized experts in the research field of porphyrins each having his own separate area of expertise in the field Between them they have published over 1500 peer reviewed papers and edited more than three dozen books on diverse topics of porphyrins and phthalocyanines In assembling the new volumes of this unique handbook they have selected and attracted the very best scientists in each sub discipline as contributing authors This handbook will prove to be a modern authoritative treatise on the subject as it is a collection of up to date works by world renowned experts in the field Complete with hundreds of figures tables and structural formulas and thousands of literature citations all researchers and graduate students in this field will find the Handbook of Porphyrin Science an essential major reference source for many years to come

Mechanisms of Inorganic and Organometallic Reactions M. V. Twigg, 1983 **Mechanisms of Inorganic and Organometallic Reactions**, 1983 *Canadian Journal of Chemistry*, 1986 *Volume 1 Elements and Inorganic Compounds* Y. S. Touloukian, J. Koolhaas Gerritsen, W. H. Shafer, 1973-01-09 The phenomenal growth of science and technology As a complementary effort to its Data Tables has brought about a universal appreciation of the fact Series TPRC published in 1967 a work entitled that present limitations in many technical develop Thermo physical Properties Research Literature Re ments are often a direct result of the paucity of knowl trieval Guide This three book work reported 33 700 edge on material properties Engineering develop references on seven thermophysical property groups ments in the years ahead will be closely linked to the and about 45 000 materials This Basic Edition sys research that is done today to contribute to a better tematically covered the world s unclassified literature understanding of the properties of matter of which published essentially between 1920 and mid 1964 in thermophysical properties constitute a major segment many instances going much earlier While research on the properties of materials con The present work referred to as Supplement I to tinues adequate steps are not being taken to ensure the Basic Edition reports an additional 26 000 refer that this invaluable body of

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