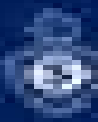
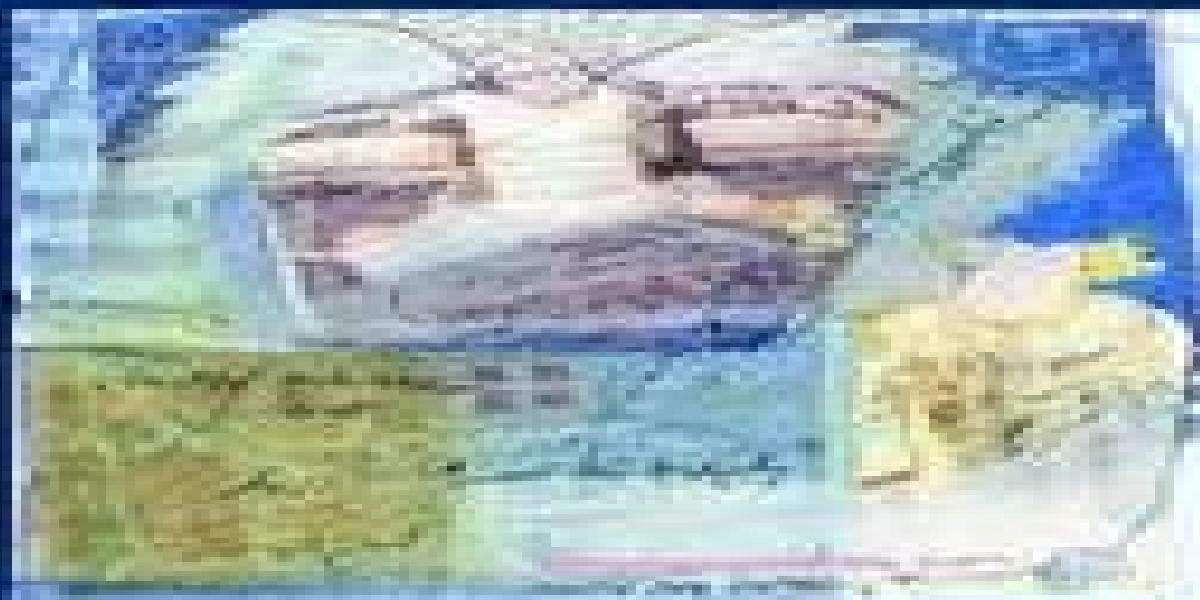


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Editors

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web log [H⁺] = 10^{-4.05} = 8.91 × 10⁻⁵ mol dm⁻³ log pK_a = pH = 3.75 + 4.05 = 7.80 and HCOOH (0.1 M) so HCOO⁻ (0.200 mol dm⁻³) amount vol x

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