

El sistema periódico

Gracias al descubrimiento de la electrolisis, se descubrieron una serie de elementos químicos. Gracias a la espectroscopia, hubo un gran avance en el descubrimiento de los elementos químicos. Actualmente se conocen 118 elementos. Una vez empezó el descubrimiento de los elementos químicos, se empezó a ver cómo podemos clasificarlos. La primera clasificación la debemos a **Lavendrier**. Ordenó los 33 "elementos" conocidos (además de la luz, el calórico y algunos compuestos) en una tabla de sustancias simples, que fue el primer esbozo de una tabla ordenada de los elementos químicos.

Un elemento químico es toda sustancia que los químicos no puedan separar en otras más simples por medios químicos.

GRUPOS	ELEMENTOS
Sin grupo	Luz
Grupo 1: Sustancias que pueden ser los elementos de los cuerpos	Calórico, oxígeno, azoe, hidrógeno
Grupo 2: Sustancias no metálicas, oxidables y acidificables	azufre, fósforo, carbono, radical muriático, radical fluorico, radical borácico
Grupo 3: Sustancias metálicas, oxidables y acidificables	antimonio, plata, arsénico, bismuto, cobalto, cobre, estaño, hierro, manganeso, mercurio, molibdeno, nickel, oro, platino, plomo, tungsteno, zinc
Grupo 4: Salificables térreas	cal, magnesia, benta, alúmina, sílica

Poco después **Dobereiner** escribió la **ley de las triadas**. Manifestó la existencia de grupos de tres elementos, triadas, que presentaban analogías en sus propiedades físicas y químicas. Ordenadas por sus masas atómicas, la masa del elemento central era aproximadamente la media aritmética de las masas de los otros dos.

ELEMENTOS	MASAS ATÓMICAS	MEDIA	PROPIEDADES
Li Na K	7 23 39	23	Metales muy reactivos
Ca Sr Ba	40 88 137	88,5	Metales reactivos
S Se Te	32 79 126	79	No metales coloreados
Cl Br I	35 80 127	81	No metales muy reactivos

Congreso de Karlsruhe

Celebrado en 1860. Fue el primer congreso internacional de Química. Se estableció claramente la diferencia entre los conceptos de átomo y molécula.

Cannizzano defendió la hipótesis de Avogadro, hasta entonces desconocida por la mayoría de los presentes. En base a dicha hipótesis, presentó una nueva tabla de pesos atómicos, que sirvió de base para el rápido desarrollo posterior de la tabla periódica.

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