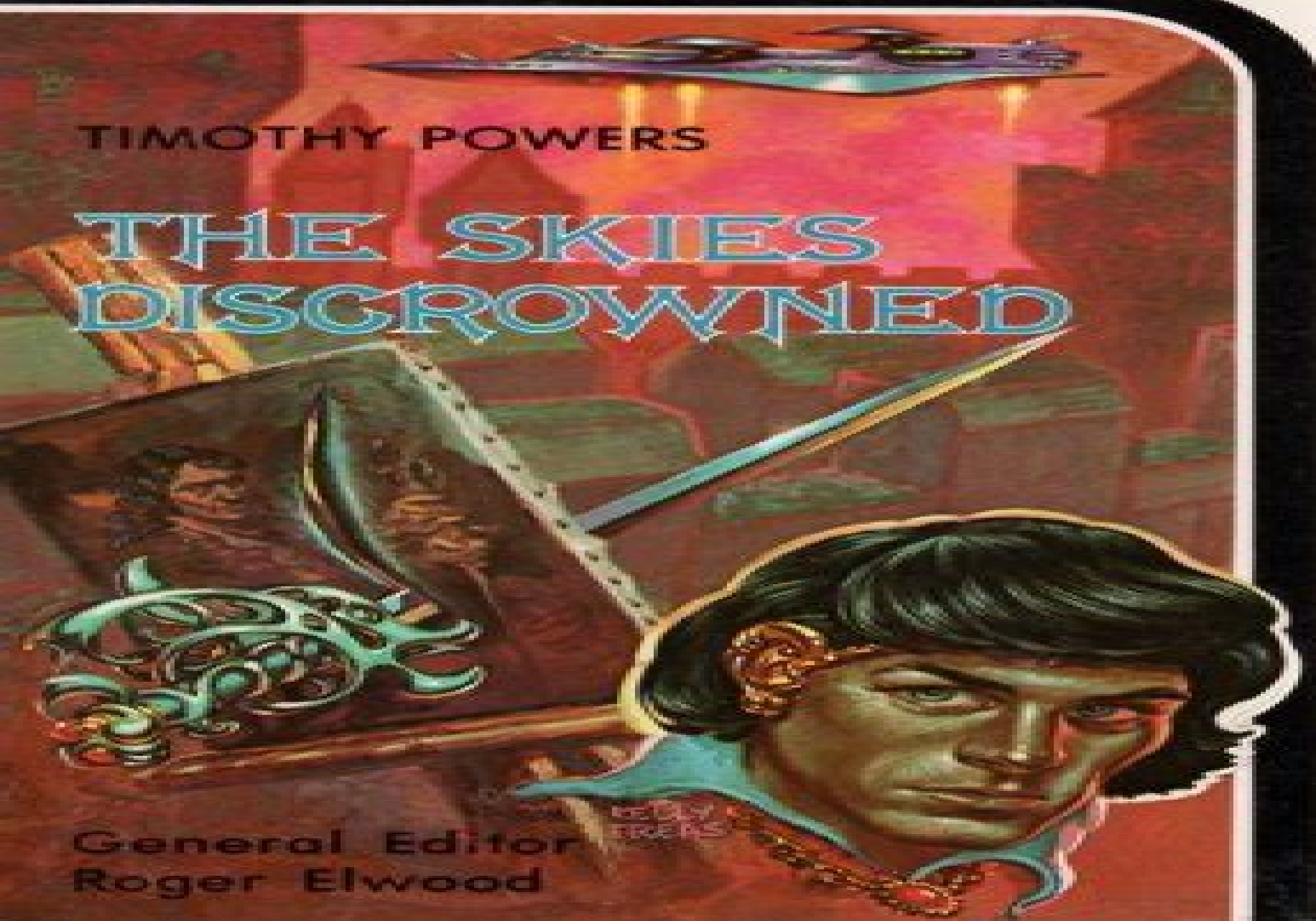


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THE SKIES DISCROWNED

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solution of molarity 1.5 M requires 1.5 mol of Na to every litre of solvent

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 where M_1, M_2, \dots, M_n are the molarities of the components and V_1, V_2, \dots, V_n are their respective volumes. For example, if you have 100 ml of 1.5 M NaCl and 200 ml of 2.5 M NaCl, the total volume is 300 ml and the total moles of NaCl are $0.15 + 0.5 = 0.65$ mol. The molarity of the mixture is $0.65 / 0.3 = 2.17$ M.

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