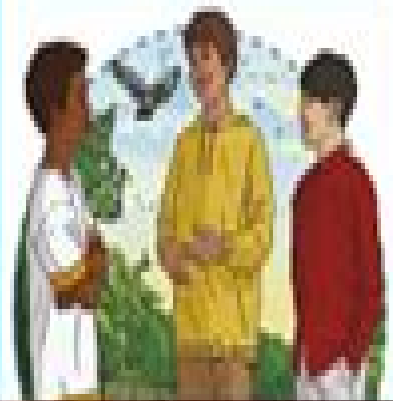
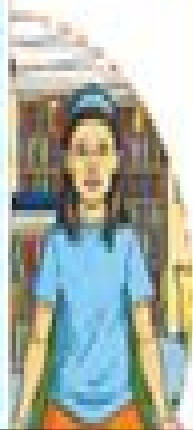


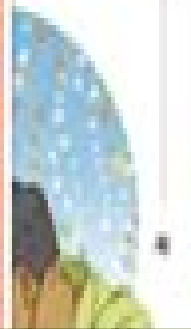
Focus on the person who is talking.



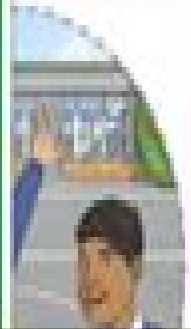
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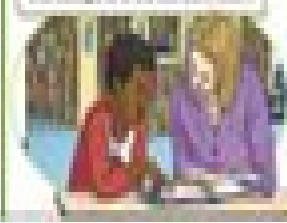
Why do you think this?

The talk is a conversation.

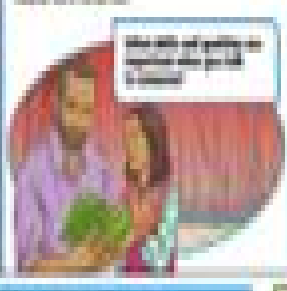


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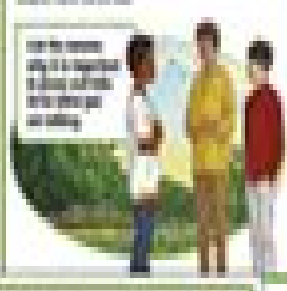
What would you do to help them understand?



Why do you think this?



Why do you think this?



Did I listen?

Listening often involves not understanding what is being said. We need to focus on the person who is talking and not be distracted by people or things that are going on around us.



When we are in a group, we often hear the speaker and the other person saying the same. This can make it hard to hear the speaker and not hear the other.

The speaker needs to know exactly what they are saying. The speaker has to listen carefully. They need to hear the other person saying the same. The speaker and the listener can check how well they have understood each other.

How can I listen better?



How Do You Know When You Are Being Listened To?

What does it mean when you feel like you are being listened to?



What did the person who was listening to you do that made you feel like this?

Remember that people who are not listening are not listening.

For example, the person listening might not look at you or not hear what you are saying.



Write your ideas in the space below.

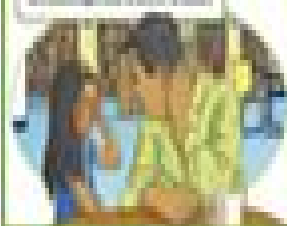
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Taking the Time to Talk and Listen



Why do you think this?

What would you do to help them understand?



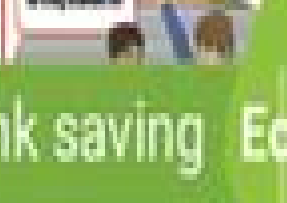
Why do you think this?

What would you do to help them understand?



Why do you think this?

What would you do to help them understand?



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Taking Time To Talk

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have $v_m = \frac{32 \text{ g mol}^{-1}}{1790 \text{ g l}^{-1}} = 0.0179 \text{ l mol}^{-1}$

web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contain 6 02 10 23 molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6×10^{23} particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6×10^{23} but as different particles have different mass their molar mass varies

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