

Physical Science Study Guide Unit 2- Atomic Structure

1) What are the three subatomic particles, and where are they located in an atom?

Answer: protons and neutrons located within the nucleus, with electrons in orbitals surrounding the nucleus

2) Which subatomic particles are used to calculate the atomic number? Which subatomic particles are used to calculate the mass number?

Answer: the number of protons in an atom, protons and the number of neutrons

3) The different rows of the periodic table represent what about the electrons of an atom?

Answer: a new outer shell of electrons

4) What is a bond? What are atoms trying to do when they bond?

Answer: to attain a stable valence electron configuration (8 valence electrons) by bonding with other atoms. by forming covalent bonds

5) Where on the periodic table does it tell you the number of valence electrons an element has?

Answer: the column number that it is in

6) The atoms on the periodic table are organized by an increasing number of what?

Answer: number of protons

7) Why are the Noble gases unreactive?

Answer: Because they're reluctant to share electrons from their filled outer electron shells

8) Where on the periodic table do you find elements that are reactive because they have high electronegativity?

Answer: top right

9) Which elements are the best conductors: Metals or nonmetals?

Answer: Copper, Nickel, Tin

10) Where are nonmetals located on the periodic table?

Answer: right

11) Why is Hydrogen's Mass Number the same as its Atomic Number?

Answer: Presence of no neutron

12) Ionic bonds are most likely to occur when the number of valence electrons is close to what numbers?

Answer: The name of the compound is sodium fluoride with formula NaF.

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Table of Contents Sms One Point Two Study Guide

1. Understanding the eBook Sms One Point Two Study Guide
 - The Rise of Digital Reading Sms One Point Two Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Sms One Point Two Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sms One Point Two Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sms One Point Two Study Guide
 - Personalized Recommendations
 - Sms One Point Two Study Guide User Reviews and Ratings
 - Sms One Point Two Study Guide and Bestseller Lists
5. Accessing Sms One Point Two Study Guide Free and Paid eBooks
 - Sms One Point Two Study Guide Public Domain eBooks

- Sms One Point Two Study Guide eBook Subscription Services
- Sms One Point Two Study Guide Budget-Friendly Options
- 6. Navigating Sms One Point Two Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Sms One Point Two Study Guide Compatibility with Devices
 - Sms One Point Two Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sms One Point Two Study Guide
 - Highlighting and Note-Taking Sms One Point Two Study Guide
 - Interactive Elements Sms One Point Two Study Guide
- 8. Staying Engaged with Sms One Point Two Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sms One Point Two Study Guide
- 9. Balancing eBooks and Physical Books Sms One Point Two Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sms One Point Two Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sms One Point Two Study Guide
 - Setting Reading Goals Sms One Point Two Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sms One Point Two Study Guide
 - Fact-Checking eBook Content of Sms One Point Two Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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