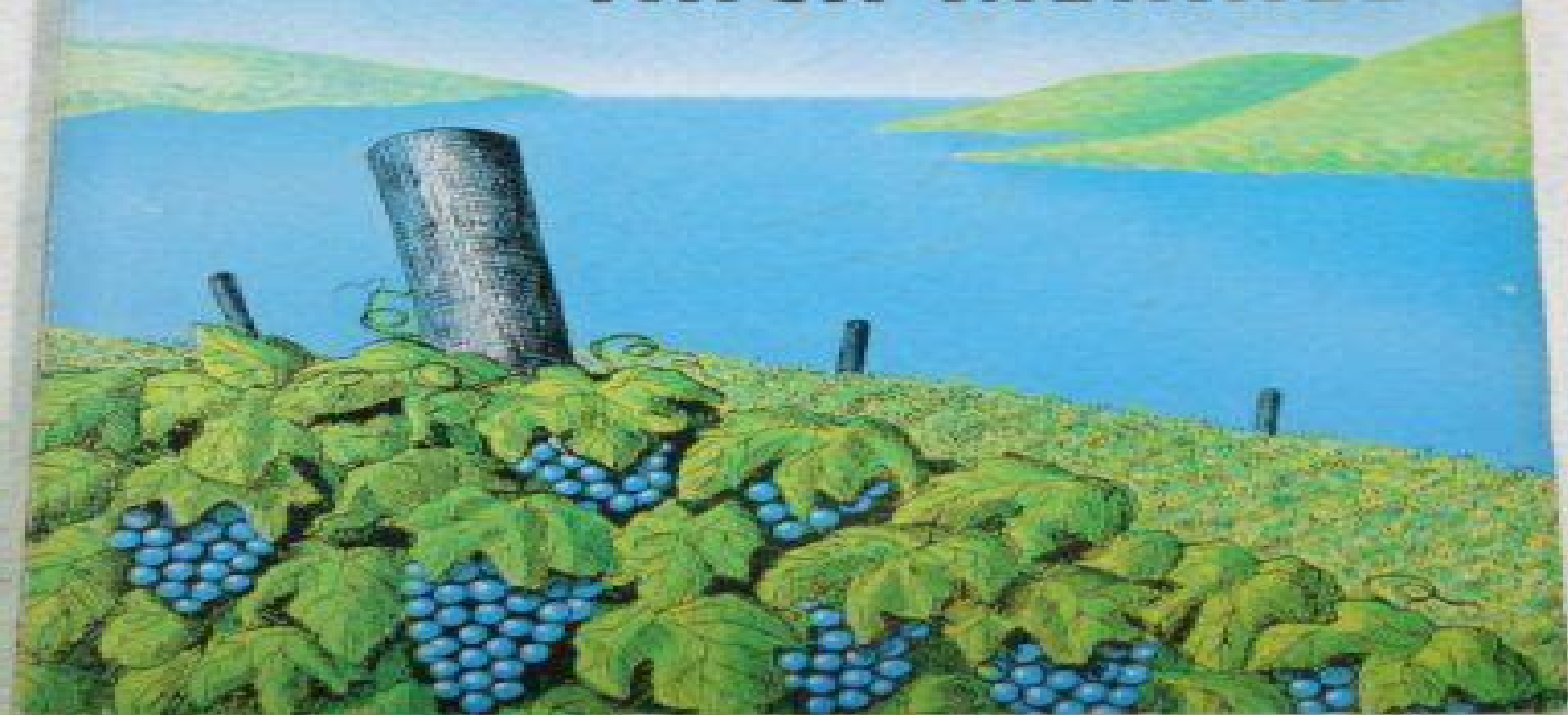


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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{t}{t_{1/2}} \times \log_2 \frac{m_0}{m}$
125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{t}{t_{1/2}} \times \log_2 \frac{m_0}{m}$
now using the half life equation in example 1 we have

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solution 20.0 / 3.6 = 5.56 half lives

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