

Henderson Hasselbalch Equation

Dissociation of an Acid



Equation

$$\text{pH} = \text{pK}_a + \log_{10} \frac{[\text{A}^{\ominus}]}{[\text{HA}]}$$

pH : pH value of the solution

K_a : Dissociation constant of the acid

$[\text{A}^{\ominus}]$: Concentration of the conjugate base (A^{\ominus})

$[\text{HA}]$: Concentration of the acid (HA)

The Henderson Equation

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