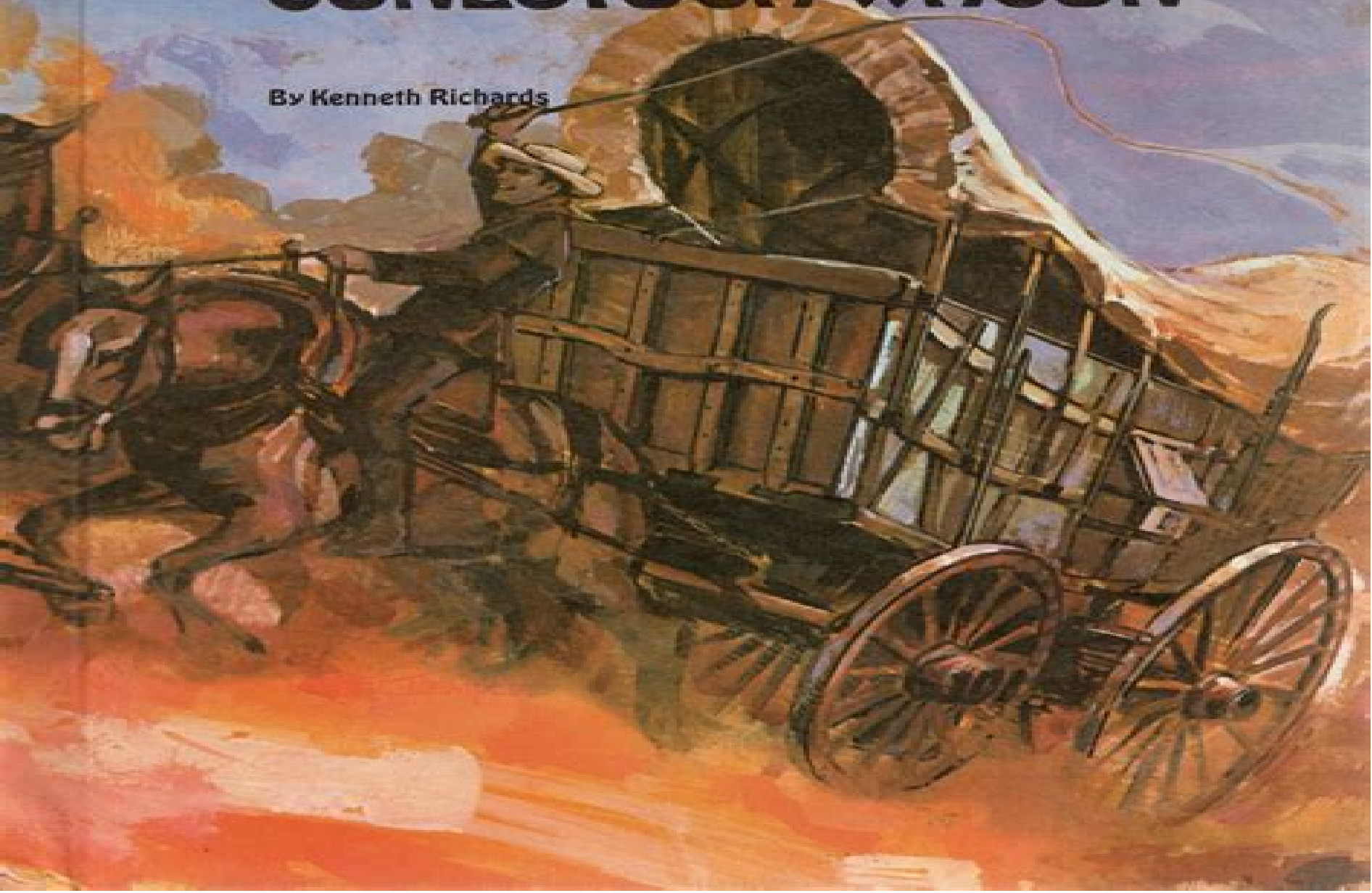


Cornerstones of Freedom

The Story of the **CONESTOGA WAGON**

By Kenneth Richards



Story Of The Conestoga Wagon

TD Snyder



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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2/\text{mole}} = 0.666 \text{ moles } \text{O}_2$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al
 $\frac{160.0 \text{ g } \text{Al}_2\text{O}_3}{101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mole}} = 1.57 \text{ moles } \text{Al}_2\text{O}_3$
 $1.57 \text{ moles } \text{Al}_2\text{O}_3 \times 26.98 \text{ g } \text{Al}_2\text{O}_3/\text{mole} = 42.4 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield
 $\frac{42.4 \text{ g } \text{Al}_2\text{O}_3}{100.26 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 42.3\%$

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web 7 Na + 2 C_2O_4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H_2 are left over 13

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