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HE STOLE THEIR HEARTS—THEN TOOK THEIR LIVES,
AND HE WOULD STRIKE AGAIN AND AGAIN—FOR HE WAS A

SMOOTH OPERATOR

THE TRUE STORY OF SEDUCTIVE SERIAL KILLER GLEN ROGERS



Sandra
486-65

Clifford L. Linedecker

Smooth Operator The True Story Of Accused Serial Killer Glen Rogers

Clifford L. Linedecker



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web in today s experiment you will determine the solubility product ksp of calcium hydroxide ca oh 2 by measuring the concentration of ca oh 2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of ca oh is represented in equation form as shown below ca oh s $\rightleftharpoons$ ca aq + 2oh aq the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution ca oh 2 $\rightarrow$ ca 2 + 2oh - 5×10^{-2} m calcium ions 5×10^{-2} m 2 0 m hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web the chemical equation used was $\text{Ca(OH)}_2 \text{ s} \rightleftharpoons \text{Ca}^{2+} \text{ aq} + 2\text{OH}^{-} \text{ aq}$ the solubility product was determined to be $K_{\text{sp}} \text{ Ca(OH)}_2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} \text{ Ca}_3(\text{PO}_4)_2$

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