

The Skaggs Mill Murders



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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 $\text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160 \text{ 0 g } \text{Al} \times 1 \text{ mol } \text{Al} / 27 \text{ g } \text{Al} \times 2 \text{ mol } \text{Al}_2\text{O}_3 / 4 \text{ mol } \text{Al} \times 101 \text{ 964 g } \text{Al}_2\text{O}_3 / 302 \text{ 3 g } \text{Al}_2\text{O}_3 = 26 \text{ 982 g } \text{Al}_2\text{O}_3$ c determine the percent yield $\text{yield actual } \times 100 = 260 \text{ 0 g } / 260 \text{ 0 g } \times 100 = 86 \text{ 01}$ theoretical 302 3 g

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web 7 $\text{Na}_2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of H_2 are left over 13

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