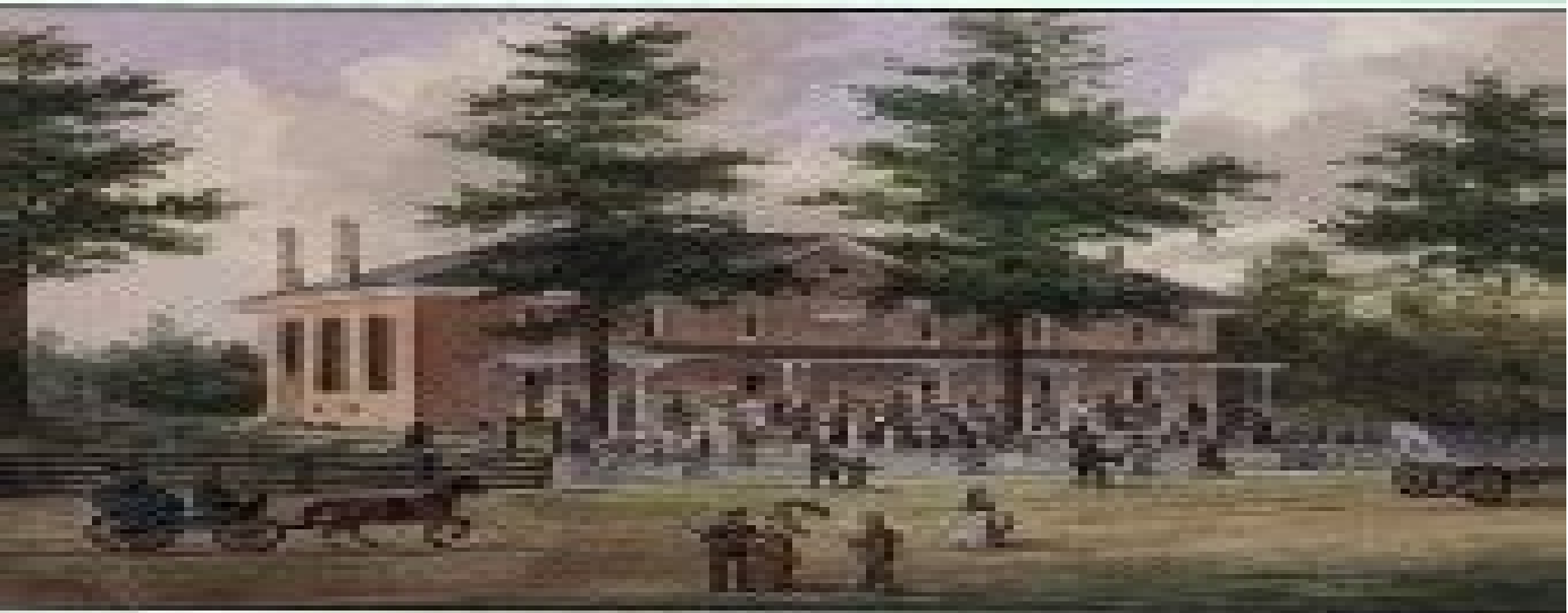


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1865-1920



Danien D. Hickey

# Sojourners No More The Quakers In The New South 186519

**Damon Douglas Hickey**



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web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

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web calorimetry worksheet 1 if 0 315 moles of hexane  $\text{C}_6\text{H}_{14}$  is combusted in a bomb calorimeter containing 5 65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55 4  $^{\circ}\text{C}$  the specific heat capacity of water is 4 184 J  $\text{g}^{-1} \text{ } ^{\circ}\text{C}^{-1}$   $m = 5 650 \text{ g}$   $\text{H}_2\text{O}$  4 184 J  $\text{g}^{-1} \text{ } ^{\circ}\text{C}^{-1}$  55 4  $^{\circ}\text{C}$   $q = 1310 \text{ kJ}$

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web 6 a 218 g sample of steam at 121  $^{\circ}\text{C}$  is cooled to ice at 14  $^{\circ}\text{C}$  find the change in heat content of the system 7 in going from ice at 34  $^{\circ}\text{C}$  to steam at 138  $^{\circ}\text{C}$  a sample of water absorbs  $1 41 \times 10^5 \text{ J}$  find the mass of the sample answers 1 2 68  $\times 10^4 \text{ J}$  2 6 23  $\times 10^4 \text{ J}$  3 1 71  $\times 10^4 \text{ J}$  4 1 11  $\times 10^5 \text{ J}$  5 2 11  $\times 10^5 \text{ J}$  6

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