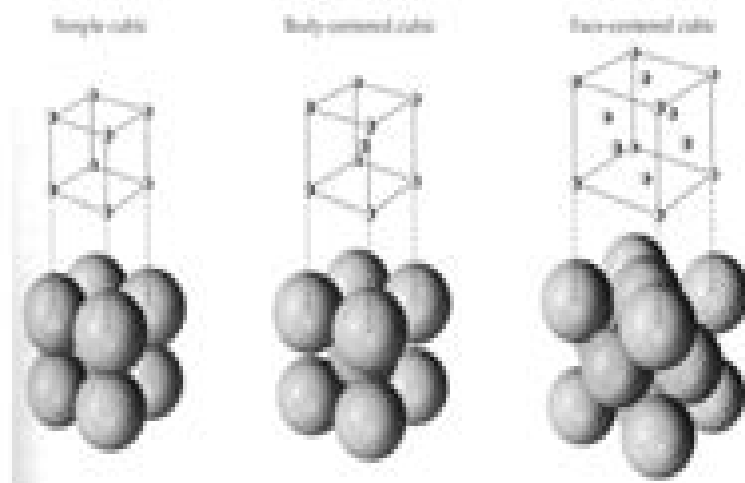


In the given figure below, simple cube is a primitive cell. Number of atoms per unit cell is one for it. Body centered cubic and face centered cubic is non primitive. No. of atoms per unit cell is 2 and 4 respectively.



Simple Cubic Structure (SC)

Simple cube has eight corner atoms. Each corner atom contribute $1/8^{\text{th}}$ to unit cell. And each atom has six nearest atom.

Number of atoms per unit cell = 8 atoms $\times 1/8 = 1$

Co-ordination number = 6



If a is lattice constant then atomic radius, $r = a/2$

Atomic packing fraction = $\frac{\text{volume of atoms}}{\text{volume of unit cell}}$

$$\begin{aligned} \text{Atomic packing fraction} &= \frac{1 \times \frac{4}{3}\pi r^3}{a^3} = \frac{\frac{4}{3}\pi r^3}{(2r)^3} = \frac{\frac{4}{3}\pi r^3}{8r^3} \\ &= \frac{\pi}{6} = 0.526 = 52.6\% \end{aligned}$$

Body Centered Cubic Structure (BCC)

Body centered cube has eight corner atoms and one atom at the center. Each corner atom contribute $1/8^{\text{th}}$ to unit cell. And center atom has eight nearest atom.

Number of atoms per unit cell = 8 atoms $\times 1/8$ (corner atom) + 1 (center atom) = 2

Co-ordination number = 8

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Lesser Blum,F.Barry Malik



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